

AP CHEM

PERIODIC TRENDS

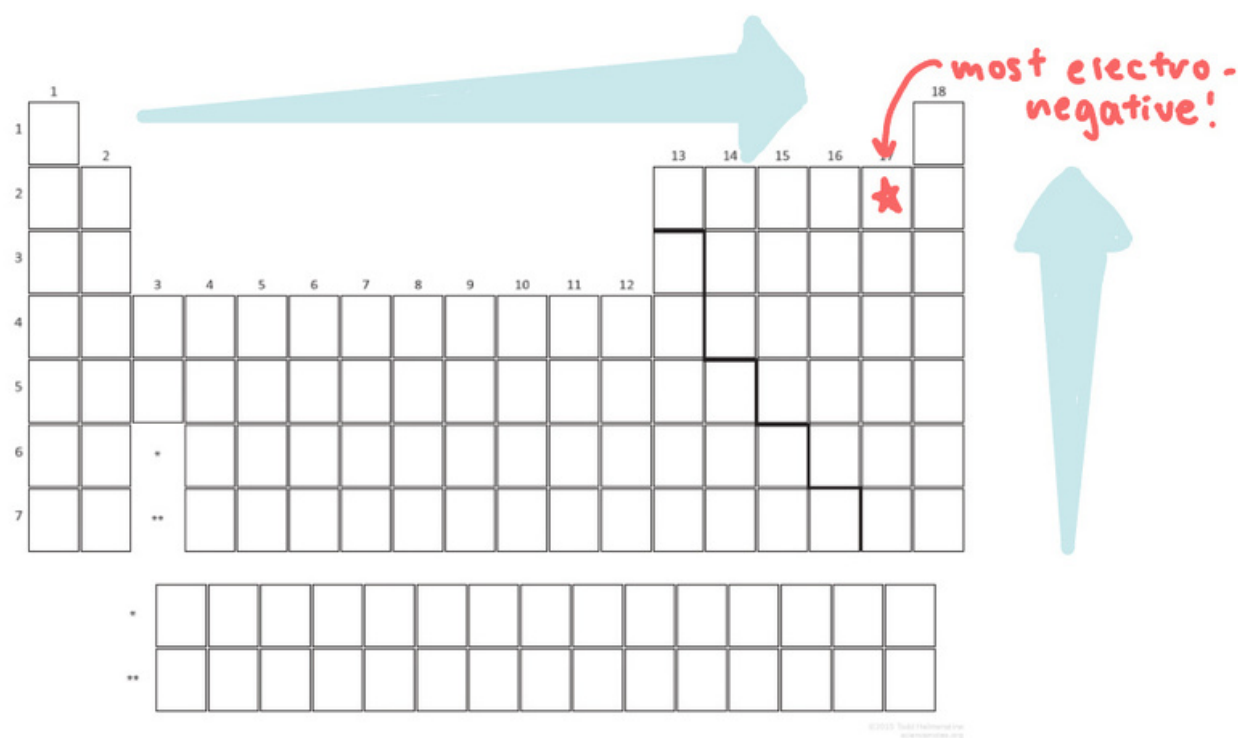
REVIEW



1. ELECTRONEGATIVITY

ELECTRONEGATIVITY: measures the tendency of an atom to attract electrons

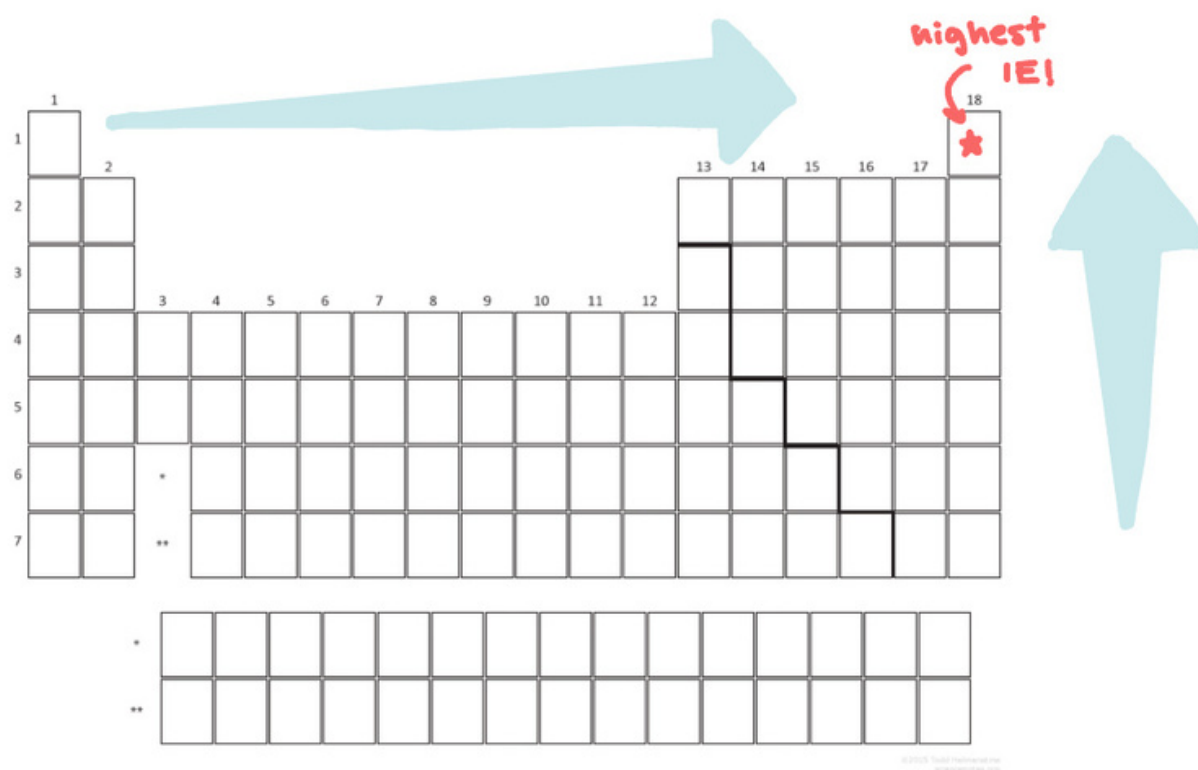
→ Think of it as how much an atom "hogs" electrons!!



★ INCREASES UPWARDS & RIGHTWARDS

2. IONIZATION ENERGY

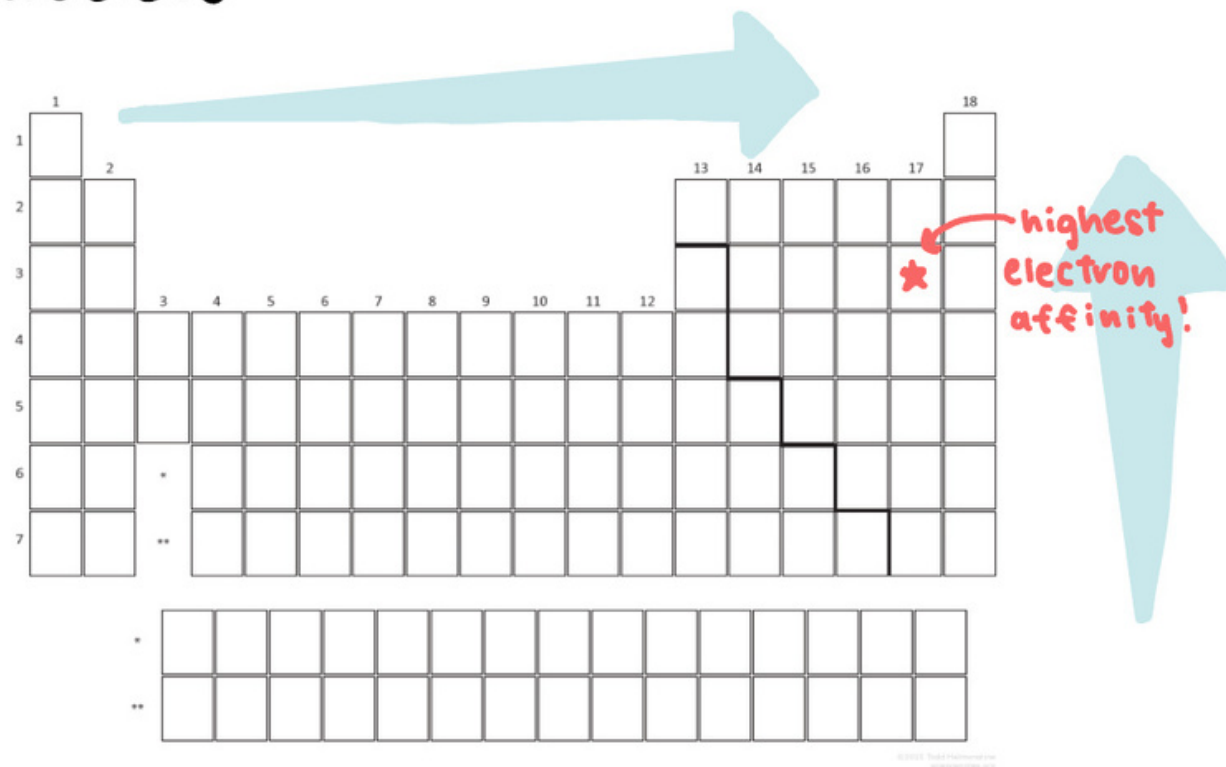
IONIZATION ENERGY : energy required to remove an outermost electron from an atom / molecule



★ INCREASES UPWARDS & RIGHTWARDS

3. ELECTRON AFFINITY

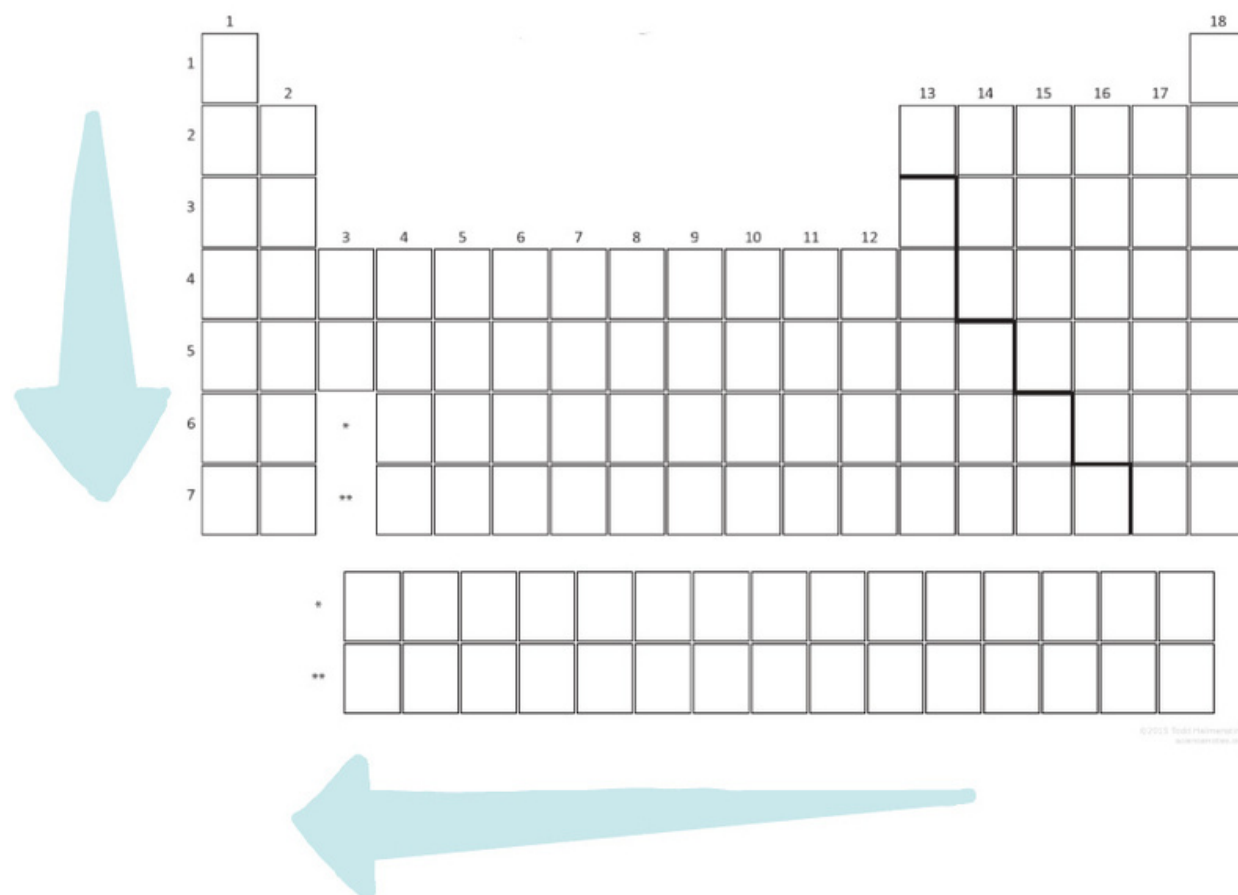
ELECTRON AFFINITY: measure of attraction between incoming electron & nucleus



*** INCREASES UPWARDS & RIGHTWARDS**

4. ATOMIC RADIUS

ATOMIC RADIUS: measure of size of an atom



* INCREASES DOWNWARDS & LEFTWARDS

5. METALLIC CHARACTER

METALLIC CHARACTER: measure of how easily an atom loses electrons (forming cations)

A periodic table diagram illustrating the trends in metallic character. The table is divided into three main sections: the s-block (groups 1 and 2), the p-block (groups 13-18), and the d-block (groups 3-10). The f-block (lanthanides and actinides) is shown below the main table. Two large, light blue arrows indicate the direction of increasing metallic character: one arrow points from the top-right (non-metallic region) towards the bottom-left (metallic region), and the other points from the top-right towards the bottom-left, labeled 'METALLIC'. The top-right region is labeled 'NON METALLIC'.

1																	18
2																	
3																	
4																	
5																	
6			*														
7			**														
			*														
			**														

*** INCREASES DOWNWARDS & LEFTWARDS**

**follow for
more!**



@mightochondria 

@mightochondria.stem 