

AP CHEM

HESS'S LAW

REVIEW



KEY IDEA:

if a reaction can be written
as a sum of several reactions,

$$\Delta H_{rxn} = \sum \Delta H_{steps}$$

• why? → enthalpy = state function,
so only initial & final states matter

MANIPULATING EQUATIONS

1. Reversing a reaction → change sign of ΔH



2. Multiply/dividing coefficients → do same to ΔH



3. Adding equations → add ΔH

HOW TO SOLVE

1. Write out target equation
2. **Compare** to given equations
3. Reverse equations (if needed)
4. Multiply coefficients to match
5. Cancel out species that are on both sides
6. **Add ΔH**

**follow for
more!**



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